

COMMON IONS

Cations

1+ Charge		2+ Charge		3+ Charge	
Formula	Name	Formula	Name	Formula	Name
H ⁺	hydrogen	Mg ²⁺	magnesium	Al ³⁺	aluminum
Na ⁺	sodium	Ca ²⁺	calcium	Fe ³⁺	iron(III)*
K ⁺	potassium	Ba ²⁺	barium		
Cu ⁺	copper(I)*	Zn ²⁺	zinc		
Ag ⁺	silver	Cd ²⁺	cadmium		
NH ₄ ⁺	ammonium	Hg ²⁺	mercury(II)*		
Li ⁺	lithium	Cu ²⁺	copper(II)*		
		Pb ²⁺	lead(II)*		
		Fe ²⁺	iron(II)*		

Anions

1- Charge		2- Charge		3- Charge	
Formula	Name	Formula	Name	Formula	Name
F ⁻	fluoride	O ²⁻	oxide	PO ₄ ³⁻	phosphate
Cl ⁻	chloride	S ²⁻	sulfide	P ³⁻	phosphide
Br ⁻	bromide	SO ₄ ²⁻	sulfate		
I ⁻	iodide	SO ₃ ²⁻	sulfite		
NO ₃ ⁻	nitrate	CO ₃ ²⁻	carbonate		
NO ₂ ⁻	nitrite				
OH ⁻	hydroxide				
HCO ₃ ⁻	hydrogen carbonate (bicarbonate)				
HOCl ⁻	hypochlorite				
SCN ⁻	thiocyanate				

*Some metal atoms form ions with one charge under certain conditions and another charge under different conditions. To specify the electrical charge for these ions, Roman numerals are used in parentheses after the metal's name.